

# Download File PDF Chemquest 45 Answer

#Jenny



Finally I get this ebook, thanks for all these I can get now!

#Rio



Cool! I'am really happy

#Markus Jensen



I did not think that this would work, my best friend showed me this website, and it does! I get my most wanted eBook

#Hun Tsu



wtf this great ebook for free?!

#Che Salsa



My friends are so mad that they do not know how I have all the high quality ebook which they do not!

#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

$2\text{N}_2\text{O}(g) \rightarrow 4\text{NO}(g) + \text{O}_2(g)$

Time (s)	Moles $\text{N}_2\text{O}$	Moles $\text{NO}$	Moles $\text{O}_2$
0	0.4720	0	0
600	0.4194	0.0252	0.0063
1200	0.4104	0.0432	0.0108
1800	0.4032	0.0576	0.0144

*Handwritten notes:*  $2(0.009)$ ,  $4(0.0036)$ ,  $0.009$ ,  $0.018$ ,  $0.0045$

**Critical Thinking Questions**

7. Calculate the change in moles of each reactant and product between 600 and 1200 seconds.

change in moles of  $\text{N}_2\text{O}$ :  $2:4$  change in moles of  $\text{NO}$ :  $4:1$  change in moles of  $\text{O}_2$ :  
 $= 0.009 + 0.018 + 0.0045$

8. Compare your answers in question 7. What is the relationship between the coefficients in the balanced chemical equation and the number of moles used up or produced in a reaction?

9. Fill in the missing blanks in the table above.

10. Calculate the following values. Be sure to use molarity in your calculations.

a) Rate of disappearance of  $\text{N}_2\text{O}$  between time 0 and 600 s.

[Download PDF version of :](#)  
**Chemquest 45 Answer**