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Chemistry 12 Unit 1 - Reaction Kinetics

19. Consider the rate of the following reaction:

$$\text{Sn(s)} + 2\text{HCl(aq)} \rightarrow \text{H}_2\text{(g)} + \text{SnCl}_2\text{(aq)}$$

a) Is it dependent on temperature? YES Explain your answer.
Mostly all reactions are temperature dependent.

b) Is it dependent on pressure? NO Explain your answer.
No gaseous reactants

c) Is it dependent on surface area? YES Explain your answer.
Heterogeneous reactions. (s) and (aq) reactants

20. Consider the following reaction:

$$2\text{NO(g)} + 2\text{H}_2\text{(g)} \rightarrow \text{N}_2\text{(g)} + 2\text{H}_2\text{O(g)}$$

Data collected for the above reaction was used to construct the following graph:

Time (s)	Molar NO
0	0.04
10	0.02

From this graph, determine the rate of reaction in moles of NO consumed per second.

Slope = $-\frac{(0.02 - 0.04) \text{ mol}}{(10 - 0) \text{ s}} = \frac{0.02 \text{ mol}}{10 \text{ s}} = 0.002 \text{ mol/NO/s}$

so rate = 0.002 mol/NO/s Answer: 0.002 mol/NO/s

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